

STUDENT LAB SHEET:



HS - PS1 - 2&7

STOICHIOMETRY

Introduction: You learned in Chapter 4 that the mole ratio is the ratio of the number of moles of reactants to the number of moles of products. In this lab you will measure the amount of product formed when iron(II) sulfate reacts with zinc. From this information, you will determine which reaction, shown below, applies to the reaction you see in this experiment. The possible reactions are:



Pre-lab Questions: What is the formula for the iron(II) sulfate? A compound of iron(II) sulfate, zinc(II) sulfate reacts with zinc. How?

Hypothesis: Should attempt to explain whether you expect your reaction will be the one the following procedure?

Pre-laboratory Questions:

- 1) Write out and balance each of the following equations:
 - a. $\text{Zinc}(s) + \text{Iron(II) sulfate (aq)} \rightarrow \text{Iron(II) sulfate (aq)} + \text{Zinc(II) sulfate (aq)}$
 - b. $\text{Zinc(II) sulfate (aq)} + \text{Iron(II) sulfate (aq)} \rightarrow \text{Iron(II) sulfate (aq)} + \text{Zinc(II) sulfate (aq)}$
- 2) What is the limiting reagent in the reaction below? Equations of: Stoichiometry
- 3) Using your limiting reactant, if 20g of iron(II) sulfate were dissolved, what mass of zinc(II) sulfate would be expected?
- 4) Using above limiting reactant, if 20g of iron(II) sulfate were dissolved, what mass of zinc(II) sulfate would be expected?

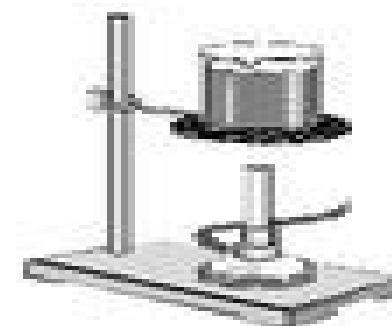
Materials:

Apparatus:

Bottle grinders
Beaker
Bunsen burner
Ring stand
Spoon
Stainless beakers (2)
Tongs
Digital scale
Eudiometer tube
Spatula
Stopper
Stopper
Hot plate with variable control

Reagents:

Chromic (II) sulfate
Water



Safety:

- 1) Do not plug/unplug electrical items in the room in this laboratory.
- 2) Wear all provided safety items in this lab. This includes lab coat, plastic apron, clothing.
- 3) If you come in contact with any solutions, wash it immediately, with cold water and dry thoroughly.

Procedure:

- 1) Weigh a clean, dry 250mL beaker. Record mass in data table provided.
- 2) Place 7.00g of zinc(II) sulfate in a beaker.
- 3) Add about 50mL of water to the beaker.
- 4) Arrange the breaker, wire grinders, Bunsen burner and ring stand as shown in the image to the right.

Experiment 37 Stoichiometry Answers

Jicheng Xie

Experiment 37 Stoichiometry Answers:

Experiment 37 Stoichiometry Answers Book Review: Unveiling the Magic of Language

In an electronic digital era where connections and knowledge reign supreme, the enchanting power of language has been much more apparent than ever. Its ability to stir emotions, provoke thought, and instigate transformation is truly remarkable. This extraordinary book, aptly titled "**Experiment 37 Stoichiometry Answers**," published by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we will delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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